



TOXICS RELEASE INVENTORY

List of Toxic Chemicals within the Water Dissociable Nitrate Compounds Category and Guidance for Reporting

Section 313 of the Emergency Planning and Community Right-to-Know Act of 1986 (EPCRA) requires certain facilities manufacturing, processing, or otherwise using listed toxic chemicals to report their environmental releases of such chemicals annually. Beginning with the 1991 reporting year, such facilities also must report pollution prevention and recycling data for such chemicals, pursuant to section 6607 of the Pollution Prevention Act, 42 U.S.C. 13106. When enacted, EPCRA section 313 established an initial list of toxic chemicals that was comprised of more than 300 chemicals and 20 chemical categories. EPCRA section 313(d) authorizes EPA to add chemicals to or delete chemicals from the list, and sets forth criteria for these actions.

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FOREWORD

This document is an updated version of the previous document, EPA 745-R-99-008, June 1999. This version has the following updates:

- In SIC code table on page 3, the SIC code for the industrial sector of coal mining has been changed from 12 (except 12411) to 12 (except 1241); and
- In Example 5 on page 9, the unit mol/liter is changed to lb/gal.

Section 1. Introduction

On November 30, 1994, EPA added 286 chemicals and chemical categories, which include 39 chemicals as part of two delineated categories, to the list of toxic chemicals subject to reporting under section 313 of the Emergency Planning and Community Right-to-Know Act of 1986 (EPCRA), 42 U.S.C. 11001. These additions are described at 59 FR 61432, and are effective January 1, 1995 for reports due July 1, 1996. Six chemical categories (nicotine and salts, strychnine and salts, polycyclic aromatic compounds, water dissociable nitrate compounds, diisocyanates, and polychlorinated alkanes) are included in these additions. At the time of the addition, EPA indicated that the Agency would develop, as appropriate, interpretations and guidance that the Agency determines are necessary to facilitate accurate reporting for these categories. This document constitutes such guidance for the water dissociable nitrate compounds category.

Section 1.1 Who Must Report

A facility is subjected to the provisions of EPCRA section 313, if it meets all three of the following criteria:

- C It is included in a covered Standard Industrial Classification (SIC) code as listed in the following table; and

Industrial Sector	SIC code
Metal mining	10 (except 1011, 1081, and 1094)
Coal Mining	12 (except 1241)
Manufacturing	20-39
Electrical utilities	4911, 4931, and 4939 (limited to facilities that combust coal and/or oil for the purpose of generating electricity for distribution in commerce)
Treatment, Storage, and Disposal facilities	4953 (limited to facilities regulated under the RCRA Subtitle C, 42 U.S.C. section 6921 <i>et seq.</i>)
Chemical distributors	5169
Petroleum bulk terminals	5171
Solvent recovery services	7389 (limited to facilities primarily engaged in solvent recovery services on a contract or fee basis)

- It has 10 or more full-time employees (or the equivalent 20,000 hours per year); and
- It manufactures, imports, processes, or otherwise uses any of the toxic chemicals listed on the EPCRA section 313 list in amounts greater than the “threshold” quantities specified below.

Section 1.2 Thresholds

Thresholds are specified amounts of toxic chemicals used during the calendar year that trigger reporting requirements.

If a facility *manufactures* or *imports* any of the listed toxic chemicals, the threshold quantity will be:

- 25,000 pounds per toxic chemical or category during the calendar year.

If a facility *processes* any of the listed toxic chemicals, the threshold quantity will be:

- 25,000 pounds per toxic chemical or category during the calendar year.

If a facility *otherwise uses* any of the listed toxic chemicals (without incorporating it into any product, or producing it at the facility), the threshold quantity is:

- 10,000 pounds per toxic chemical or category during the calendar year.

EPCRA section 313 requires threshold determinations for chemical categories to be based on the total of all chemicals in the category manufactured, processed, or otherwise used. For example, a facility that manufactures three members of a chemical category would count the total amount of all three chemicals manufactured towards the manufacturing threshold for that category. When filing reports for chemical categories, the releases are determined in the same manner as the thresholds. One report is filed for the category and all releases are reported on this form.

Section 1.3 Chemicals within the Water Dissociable Nitrate Compounds Category

EPA is providing a list of CAS numbers and chemical names to aid the regulated community in determining whether they need to report for the water dissociable nitrate compounds category. The list includes individual chemicals within the water dissociable nitrate compounds category. If a facility is manufacturing, processing, or otherwise using a chemical which is on this list, they must report this chemical. However, this list is not exhaustive. If a facility is manufacturing, processing, or otherwise using a water dissociable nitrate compound, they must report the chemical, even if it does not appear on the list.

Section 1.4 *De Minimis* Concentrations

The water dissociable nitrate compounds category is subject to the one percent *de minimis* concentration. Thus, mixtures that contain members of this category equal to or in excess of the *de minimis* should be factored into threshold and release determinations.

Section 2. Guidance for Reporting Chemicals within the Water Dissociable Nitrate Compounds Category

Note: for the purposes of reporting under the nitrate compounds category, water dissociable means that the nitrate ion dissociates from its counterion when in solution.

Section 2.1 Chemicals within the Water Dissociable Nitrate Compounds Category

Chemicals within the nitrate compounds category are only reportable when in aqueous solution. All water dissociable nitrate compounds are included in the nitrate compounds category, including ammonium nitrate. Specifically listed section 313 chemicals *are not* included in threshold determinations for chemical categories such as the water dissociable nitrate compounds category. Specifically listed toxic chemicals are subject to their own individual threshold determinations. As of December 1, 1994, ammonium nitrate (solution) is not an individually listed chemical on the EPCRA section 313 list. However, ammonium nitrate is still subject to reporting under the nitrate compounds category. In addition, the aqueous ammonia from the dissociation of ammonium nitrate when in aqueous solution is subject to reporting under the ammonia listing.

Section 2.2 Determining Threshold and Release Quantities for Nitrate Compounds

The total nitrate compound, including both the nitrate ion portion and the counterion, is included in the nitrate compounds category. When determining threshold amounts, the total weight of the nitrate compound is to be included in all calculations. However, only the nitrate ion portion is to be included when determining the amount of the chemicals within the nitrate compounds category that is released, transferred, or otherwise managed in wastes.

Example 1: In a calendar year, a facility processes 100,000 pounds of ammonium nitrate (NH_4NO_3), in aqueous solution, which is released to wastewater streams, then transferred to a POTW. The quantity applied towards threshold calculations for the nitrate compounds category is the total quantity of the nitrate compound or 100,000 pounds. Since this quantity exceeds the 25,000 pound processing threshold, the facility is required to report for the nitrate compounds category. Under the nitrate compounds category, only the weight of the nitrate ion portion of ammonium nitrate is included in release and transfer calculations. The molecular weight of ammonium nitrate is 80.04 and the weight of the nitrate ion portion is 62.01 or 77.47 percent of the molecular weight of ammonium nitrate. Therefore, the amount of nitrate ion reported as transferred to the POTW is 77.47 percent of 100,000 pounds, or 77,470 pounds (reported as 77,000 pounds). The aqueous ammonia from ammonium nitrate is reportable under the EPCRA Section 313 listing for ammonia. For determining thresholds and calculating releases under the ammonia listing, see the separate directive, *Guidance for Reporting Aqueous Ammonia* (EPA document #745-R-00-004, Revised December, 2000).

Example 2: In a calendar year, a facility manufactures as by-products 20,000 pounds of sodium nitrate (NaNO_3) and 10,000 pounds of calcium nitrate ($\text{Ca}(\text{NO}_3)_2$), both *in aqueous solutions*, and releases these solutions to wastewater streams. The total quantity of nitrate compounds manufactured by the facility is the sum of the two chemicals, or 30,000 pounds, which exceeds the manufacturing threshold quantity of 25,000 pounds. The facility therefore is required to report for the nitrate compounds category. By weight, the nitrate ion portion is 72.96 percent of sodium nitrate and is 75.57 percent of calcium nitrate. Of the 20,000 pounds of the sodium nitrate in solution, 72.96 percent or 14,592 pounds is nitrate ion, and similarly, of the 10,000 pounds of the calcium nitrate in solution, 75.57 percent or 7,557 pounds is nitrate ion. The total nitrate ion in aqueous solution released by the facility is the sum of the nitrate ion in the two solutions, or 22,149 pounds.

Section 2.3 Reporting Nitrate Compounds Generated from the Partial or Complete Neutralization of Nitric Acid

Nitric acid is an individually listed chemical on the original EPCRA section 313 list and is reported as a separate chemical if the manufacture, process or otherwise use thresholds are exceeded. The partial or complete neutralization of nitric acid results in the formation of nitrate compounds which are reported as chemicals within the nitrate compounds category if their manufacture, process or otherwise use thresholds are exceeded.

Mineral acids such as nitric acid may be present in aqueous waste streams that are sent to on-site neutralization or are discharged to a publicly owned treatment works (POTW) or other off-site treatment facility. As stated in the current version of *Toxic Chemical Release Inventory Reporting Form R and Instructions* document, on-site acid neutralization and its efficiency must be reported in Part II, section 7A of Form R (waste treatment methods and efficiency section). For purposes of reporting on Form R, EPA considers a waste mineral acid at a pH 6.0 or higher to be 100 percent neutralized (water discharges to receiving streams or POTWs are reported as zero). The nitrate compounds produced from the complete neutralization (pH 6.0 or above) of nitric acid are reportable under the nitrate compounds category and should be included in all threshold and release calculations. Two Form R reports would be required if the manufacture, process or otherwise use thresholds are exceeded for nitric acid and for the nitrate compounds category.

If the nitric acid treatment efficiency is not equal to 100 percent (pH is less than 6.0), the amount of the acid remaining in the waste stream which is released to the environment on-site or off-site must be reported in Part II of Form R. The nitrate compounds produced from the partial neutralization of nitric acid are reportable under the nitrate compounds category and should be included in all threshold and release calculations. Two reports would again be required if the manufacture, process or otherwise use thresholds are exceeded for nitric acid and for the nitrate compounds category.

Section 2.3.1 Estimating Nitric Acid Releases

The pH of the waste stream can be used to calculate the amount of nitric acid in the stream and the efficiency of neutralization. The pH is a measure of the acidity or alkalinity of a waste stream and can be obtained readily using a pH meter or pH sensitive paper. The pH scale itself varies from 0.0 to 14.0.

The total nitric acid concentration (ionized and unionized) in pounds/gallon can be calculated by using the pH value of the solution, the molecular weight and ionization constant of the acid, and appropriate conversion factors. The total acid concentration for nitric acid for different pH values is listed in Table 1. The calculation of mineral acid concentrations and the derivation of Table 1 are discussed in a separate directive, *Estimating Releases for Mineral Acid Discharges Using pH Measurements*, and in an addendum to this directive.

The procedure outlined in this guidance document for calculating the quantity of nitrate compounds formed from the complete or partial neutralization of nitric acid can be used if nitric acid is the only mineral acid in a solution. In addition, the calculation of nitric acid releases using only pH measurements is a rough estimate. The subsequent calculation of nitrate compound releases is therefore also only a rough estimate. The estimates can be made for a waste stream with a steady pH below 6.0 or for one whose pH temporarily drops to below pH 6.0. Facilities should use their best engineering judgement and knowledge of the solution to evaluate how reasonable the estimates are.

Example 3: In a calendar year, a facility transfers 1.0 million gallons of a solution containing nitric acid (HNO₃), at pH 4.0, to a POTW. Using Table 1 (next page), a pH of 4.0 corresponds to a concentration of 0.0000520 lb HNO₃/gallon of solution. The weight of HNO₃ transferred can be estimated using the equation:

$$\text{Transfer of HNO}_3 = (\text{concentration of HNO}_3) \times (\text{effluent flow rate})$$

Substituting the example values into the above equation yields:

$$\text{Transfer of HNO}_3 = 0.0000520 \text{ lb/gal HNO}_3 \times 1,000,000 \text{ gal solution/yr} = 52 \text{ lb/yr}$$

Example 4: A facility had an episodic release of nitric acid (HNO₃) in which the waste stream was temporarily below pH 6.0. During the episode, the waste water (pH 2.0) was discharged to a river for 20 minutes at a rate of 100 gallons per minute. Using Table 1, a pH of 2.0 for HNO₃ represents a concentration of 0.0052000 lb HNO₃/gallon of solution. The amount of the HNO₃ released can be estimated using the following equation:

$$\text{Release of HNO}_3 = (\text{concentration of HNO}_3) \times (\text{effluent flow rate}).$$

Substituting the example values in the above equation:

$$\text{Release of HNO}_3 = 0.0052000 \text{ lb/gal} \times 100 \text{ gal/min} \times 20 \text{ min} = 10 \text{ pounds}$$

Example 5: A nitric acid (HNO₃) waste-stream of pH 2.4 is neutralized to pH 4.6. Using Table 1, the initial nitric acid concentrations 0.0021000 lb/gal and the final concentration is 0.0000130 lb/gal. Substituting these values into the equation for treatment efficiency:

$$\text{Treatment Efficiency} = (0.0021000 - 0.0000130)/0.0021000 \times 100 = 99.4 \text{ percent.}$$

Table 1. Nitric Acid Concentration Versus pH

pH	Nitric Acid Concentration (lb/gal)	pH	Nitric Acid Concentration (lb/gal)
0.0	0.5200000	3.0	0.0005200
0.2	0.3300000	3.2	0.0003300
0.4	0.2100000	3.4	0.0002100
0.6	0.1300000	3.6	0.0001300
0.8	0.0830000	3.8	0.0000830
1.0	0.0520000	4.0	0.0000520
1.2	0.0330000	4.2	0.0000330
1.4	0.0210000	4.4	0.0000210
1.6	0.0130000	4.6	0.0000130
1.8	0.0083000	4.8	0.0000083
2.0	0.0052000	5.0	0.0000052
2.2	0.0033000	5.2	0.0000033
2.4	0.0021000	5.4	0.0000021
2.6	0.0013000	5.6	0.0000013
2.8	0.0008300	5.8	0.0000008
		6.0	0.0000005

Section 2.3.2 Estimating Treatment Efficiencies for Nitric Acid Neutralization

Nitric acid solutions that are neutralized to a pH of 6.0 or above have a treatment efficiency of 100 percent. If nitric acid is neutralized to a pH less than 6.0, then the reportable treatment efficiency is somewhere between 0 and 100 percent. It is possible to estimate the neutralization treatment efficiency using nitric acid concentration values directly from Table 1 in the equation given below. The concentrations correspond to the pH values before and after treatment.

$$\text{Treatment Efficiency} = (I-E)/I \times 100$$

where I = acid concentration before treatment

E = acid concentration after treatment

For strong acids only, including nitric acid, the net difference in pH before and after treatment can be used to estimate the treatment efficiency, since pH is directly proportional to the acid concentration. For example, a pH change of one unit results in a treatment efficiency of 90 percent, whether the pH change is from pH 1.0 to pH 2.0, or from pH 4.0 to pH 5.0. Table 2 summarizes treatment efficiencies for various pH changes (the pH change is the difference between the initial pH and the pH after neutralization). In Table 2, some pH changes result in the same treatment efficiency values due to rounding to one decimal place.

Table 2. Nitric Acid Treatment Efficiencies for Various pH Changes			
pH Change	Treatment Efficiency (%)	pH Change	Treatment Efficiency (%)
1.0	90.0	2.0	99.0
1.1	92.1	2.1	99.2
1.2	93.7	2.2	99.4
1.3	95.0	2.3	99.5
1.4	96.0	2.4	99.6
1.5	96.8	2.5	99.7
1.6	97.5	2.6	99.8
1.7	98.0	2.7	99.8
1.8	98.4	2.8	99.8
1.9	98.7	2.9	99.9
		3.0	99.9

Example 6: If a nitric acid (HNO_3) waste stream of pH 2.0 is treated to pH 4.0, the pH change is 2 units. Using Table 2 above, the treatment efficiency is given as 99.0 percent.

Section 2.3.3 Estimating Releases of Nitrate Compounds Generated from the Neutralization of Nitric Acid

The nitrate compounds produced from the complete neutralization (pH 6.0 or above) or partial neutralization (pH less than 6.0) of nitric acid are reportable under the nitrate compounds category if the appropriate threshold is met and should be included in all threshold and release calculations. In order to determine the quantity of a nitrate compound generated and released, the quantity of nitric acid released must be known (or calculated from the equations used in Examples 3 and 4 above) as well as the nitric acid treatment efficiency (calculated from the equations used in Examples 5 and 6 above).

The neutralization of nitric acid will most likely result in the generation of monovalent nitrate compounds (such as sodium nitrate and potassium nitrate). The quantity of these compounds formed in kilomoles will be equal to the quantity of the nitric acid neutralized in kilomoles. If divalent nitrate compounds are formed (such as calcium nitrate) the quantity of these compounds formed in kilomoles will be equal to one-half the quantity of the nitric acid neutralized in kilomoles. Similarly, if trivalent nitrate compounds are formed (such as iron (III) nitrate) the quantity formed of these compounds in kilomoles will be equal to one-third the quantity of the nitric acid neutralized in kilomoles. *Note:* To calculate the releases of nitrate compounds generated from the neutralization of nitric acid, the nitrate portion of molecular weight of the nitrate compound formed must be used. Molecular weights of some of the individual chemicals within the water dissociable nitrate compounds category are given in Table 3.

Example 7: In a calendar year, a facility transfers 50,000 pounds of nitric acid (HNO₃) to an on-site treatment facility. The nitric acid treatment efficiency is 95 percent, and the nitrate compound formed as a result of the treatment is sodium nitrate (NaNO₃). The quantity of nitric acid transferred that is neutralized (generating sodium nitrate) is 95 percent of 50,000 pounds, or 47,500 pounds. The molecular weight of nitric acid is 63.01 kg/kmol, and the molecular weight of sodium nitrate is 84.99 kg/kmol. The quantity of nitric acid neutralized is converted first to kilograms then to kilomoles using the following equations:

$$\begin{aligned}\text{Kilograms HNO}_3 \text{ neutralized} &= (\text{lb HNO}_3 \text{ neutralized}) \times (0.4536 \text{ kg/lb}) \\ \text{Kilomoles HNO}_3 \text{ neutralized} &= (\text{kg HNO}_3) \div (\text{MW of HNO}_3 \text{ in kg/kmol})\end{aligned}$$

Substituting the example values into the above equation yields:

$$\begin{aligned}\text{Kilograms HNO}_3 \text{ neutralized} &= 47,500 \text{ lb} \times 0.4536 \text{ kg/lb} = 21,546 \text{ kg} \\ \text{Kilomoles HNO}_3 \text{ neutralized} &= 21,546 \text{ kg} \div 63.01 \text{ kg/kmol} = 341.9 \text{ kmol}\end{aligned}$$

The quantity of sodium nitrate generated in kilomoles is equal to the quantity of nitric acid neutralized (341.9 kmol). The quantity of sodium nitrate generated in kilomoles is converted first to kilograms then to pounds using the following equations:

$$\begin{aligned}\text{Kilograms NaNO}_3 \text{ generated} &= (\text{kmol NaNO}_3) \times (\text{MW of NaNO}_3 \text{ in kg/kmol}) \\ \text{Pounds NaNO}_3 \text{ generated} &= (\text{kg NaNO}_3) \times (2.205 \text{ lb/kg})\end{aligned}$$

Substituting the values into the above equation yields:

$$\begin{aligned}\text{Kilograms NaNO}_3 \text{ generated} &= 341.9 \text{ kmol} \times 84.99 \text{ kg/kmol} = 29,058 \text{ kg} \\ \text{Pounds NaNO}_3 \text{ generated} &= 29,058 \text{ kg} \times 2.205 \text{ lb/kg} = 64,073 \text{ pounds (reported as 64,000 pounds)}.\end{aligned}$$

The 64,000 pounds of sodium nitrate generated is the quantity used to determine whether thresholds have been met or exceeded. The quantity of nitrate ion released is calculated as in Example 1 above.

Section 2.4 Generation of Nitrate Compounds from Biological Wastewater Treatment

If a facility treats wastewater on-site biologically, using the activated sludge process, for example, the facility may be generating nitrate compounds as by-products of this biological process. The nitrate ion generated from this process will be associated with various counterions (e.g. sodium ion, potassium ion). In the absence of information on the identity of the counterion, a facility should assume for the purposes of EPCRA section 313 threshold determinations that the counterion is sodium ion.

Section 3. CAS Number List of Some of the Individual Chemicals within the Water Dissociable Nitrate Compounds Category

EPA is providing the following list of CAS numbers and chemical names to aid the regulated community in determining whether they need to report for the water dissociable nitrate compounds category. If a facility is manufacturing, processing, or otherwise using a chemical which is listed below, they must report this chemical. However, this list is not exhaustive. If a facility is manufacturing, processing, or otherwise using a water dissociable nitrate compound, they must report this chemical, even if it does not appear on the following list.

Table 3. Listing by CAS Number of Some of the Individual Chemicals within the Water Dissociable Nitrate Compounds Category		
Chemical Name	Molecular Weight*	CAS Number
Aluminum nitrate, nonahydrate	213.00	7784-27-2
Ammonium nitrate	80.04	6484-52-2
Cerium (III) ammonium nitrate, tetrahydrate	486.22	13083-04-0
Cerium (IV) ammonium nitrate	548.23	10139-51-2
Barium nitrate	261.34	10022-31-8
Beryllium nitrate, trihydrate	133.02	7787-55-5
Cadmium nitrate	236.42	10325-94-7
Cadmium nitrate, tetrahydrate	236.42	10022-68-1
Calcium nitrate	164.09	10124-37-5
Calcium nitrate, tetrahydrate	164.09	13477-34-4
Cerium (III) nitrate, hexahydrate	326.13	10294-41-4
Cesium nitrate	194.91	7789-18-6
Chromium (III) nitrate, nonahydrate	238.01	7789-02-8
Cobalt (II) nitrate, hexahydrate	182.94	10026-22-9
*For hydrated compounds, e.g. aluminum nitrate, nonahydrate, the molecular weight <u>excludes</u> the weight of the hydrate portion. For example, the same molecular weight is provided for aluminum nitrate, nonahydrate and aluminum nitrate.		

Table 3. Listing by CAS Number of Some of the Individual Chemicals within the Water Dissociable Nitrate Compounds Category *(continued)*

Chemical Name	Molecular Weight*	CAS Number
Copper (II) nitrate, trihydrate	187.56	10031-43-3
Copper (II) nitrate, hexahydrate	187.56	13478-38-1
Dysprosium (III) nitrate, pentahydrate	348.51	10031-49-9
Erbium (III) nitrate, pentahydrate	353.27	10031-51-3
Gadolinium (III) nitrate, hexahydrate	343.26	19598-90-4
Gallium nitrate, hydrate	255.73	69365-72-6
Iron (III) nitrate, hexahydrate	241.86	13476-08-9
Iron (III) nitrate, nonahydrate	241.86	7782-61-8
Lanthanum (III) nitrate, hexahydrate	324.92	10277-43-7
Lead (II) nitrate	331.21	10099-74-8
Lithium nitrate	68.95	7790-69-4
Lithium nitrate, trihydrate	68.95	13453-76-4
Magnesium nitrate, dihydrate	148.31	15750-45-5
Magnesium nitrate, hexahydrate	148.31	13446-18-9
Manganese (II) nitrate, tetrahydrate	178.95	20694-39-7
Neodymium (III) nitrate, hexahydrate	330.25	16454-60-7
Nickel (II) nitrate, hexahydrate	182.70	13478-00-7
Potassium nitrate	101.10	7757-79-1
Rhodium (III) nitrate, dihydrate	288.92	13465-43-5
Rubidium nitrate	147.47	13126-12-0
Samarium (III) nitrate, hexahydrate	336.37	13759-83-6

*For hydrated compounds, e.g. aluminum nitrate, nonahydrate, the molecular weight excludes the weight of the hydrate portion. For example, the same molecular weight is provided for aluminum nitrate, nonahydrate and aluminum nitrate.

Table 3. Listing by CAS Number of Some of the Individual Chemicals within the Water Dissociable Nitrate Compounds Category *(continued)*

Chemical Name	Molecular Weight*	CAS Number
Samarium (III) nitrate	230.97	13465-60-6
Samarium (III) nitrate, tetrahydrate	230.97	16999-44-3
Silver nitrate	169.87	7761-88-8
Sodium nitrate	84.99	7631-99-4
Strontium nitrate	211.63	10042-76-9
Strontium nitrate, tetrahydrate	211.63	13470-05-8
Terbium (III) nitrate, hexahydrate	344.94	13451-19-9
Thorium (IV) nitrate	480.06	13823-29-5
Thorium (IV) nitrate, tetrahydrate	480.06	13470-07-0
Yttrium (III) nitrate, hexahydrate	274.92	13494-98-9
Yttrium (III) nitrate, tetrahydrate	274.92	13773-69-8
Zinc nitrate, trihydrate	189.39	131446-84-9
Zinc nitrate, hexahydrate	189.39	10196-18-6
Zirconium (IV) nitrate, pentahydrate	339.24	13986-27-1

*For hydrated compounds, e.g. aluminum nitrate, nonahydrate, the molecular weight excludes the weight of the hydrate portion. For example, the same molecular weight is provided for aluminum nitrate, nonahydrate and aluminum nitrate.